Lab 3

Determination the Concentration of Sodium Hydroxide (NaOH)

❖ **Principle :-**

$NaOH + HCl \longrightarrow NaCl + H_2O$

❖ Procedure :-

1. Fill up the burette with standard HCl (0.1N)
2. Tale 10ml of unknown NaOH solution and transfer it to clean conical flask
3. Add 2-3 drops of phenolphthalein (ph.ph) indicator
4. Titrant against standard HCl until end point at which color of phenolphthalein indicator change from pick color to colorless
5. Report the titration three times and take the average .

$V_{Initial}$	V_{Final}	Consumed Volume

❖ **Calculation**

$N_1 * V_1 = N_2 V_2$

(Analyte) (titrant)

? * 10ml =0.1*V

$N_1 = \frac{0.1 * v}{10}$ = ……

Strengh = N_1*eq.wt (NaOH) =……..g/L